

# Place O<sub>2</sub> problem

oxygen in air O<sub>2</sub> at 0°C 0.14  
 O<sub>2</sub> " 0.18

solubility in H<sub>2</sub>O at 25°C

Nitrogen 0.0143 cc/cc

Solar constant at normal incidence  
 1.92 small cal/cm<sup>2</sup> min  
 ~ 2 cal/cm<sup>2</sup> sec

O<sub>2</sub> 0.028 cc/cc

~ 8 watt/cm<sup>2</sup> ~~small~~  
 $\frac{8 \times 10^4}{60}$  watt/m<sup>2</sup>

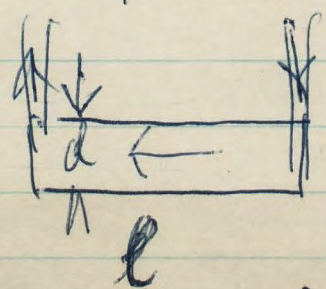
CO<sub>2</sub> 0.76 cc/cc

= 1.33 KW/m<sup>2</sup>

Carbon dioxide in air  $\frac{3}{100}$  %

if we want to grow 1 gm of dry substance  
 carbon/leaf per day and leave layer  
 10 cm thick we have to admit 1 gm  
 of carbon per 100 cm<sup>2</sup> per 24 hours  
 or  $\frac{1}{2400}$  gm per cm<sup>2</sup> per hour

or  $\frac{1}{3600} \times \frac{1}{24} \times \frac{1}{10}$  gm per cm<sup>2</sup> per sec



in 7 d sec we get  
 in 1 cm<sup>2</sup>  $\frac{1}{2}$  of CO<sub>2</sub>  
 contained in 1 l of  
 air of d l CO<sub>2</sub> contained in  
 1 cc of air.

$v = \frac{l}{7d^2}$  1 cc of CO<sub>2</sub> =  $\frac{1}{2} \times \frac{v d}{l}$

[7] = sec/air



serine registration (all?)  
numbers

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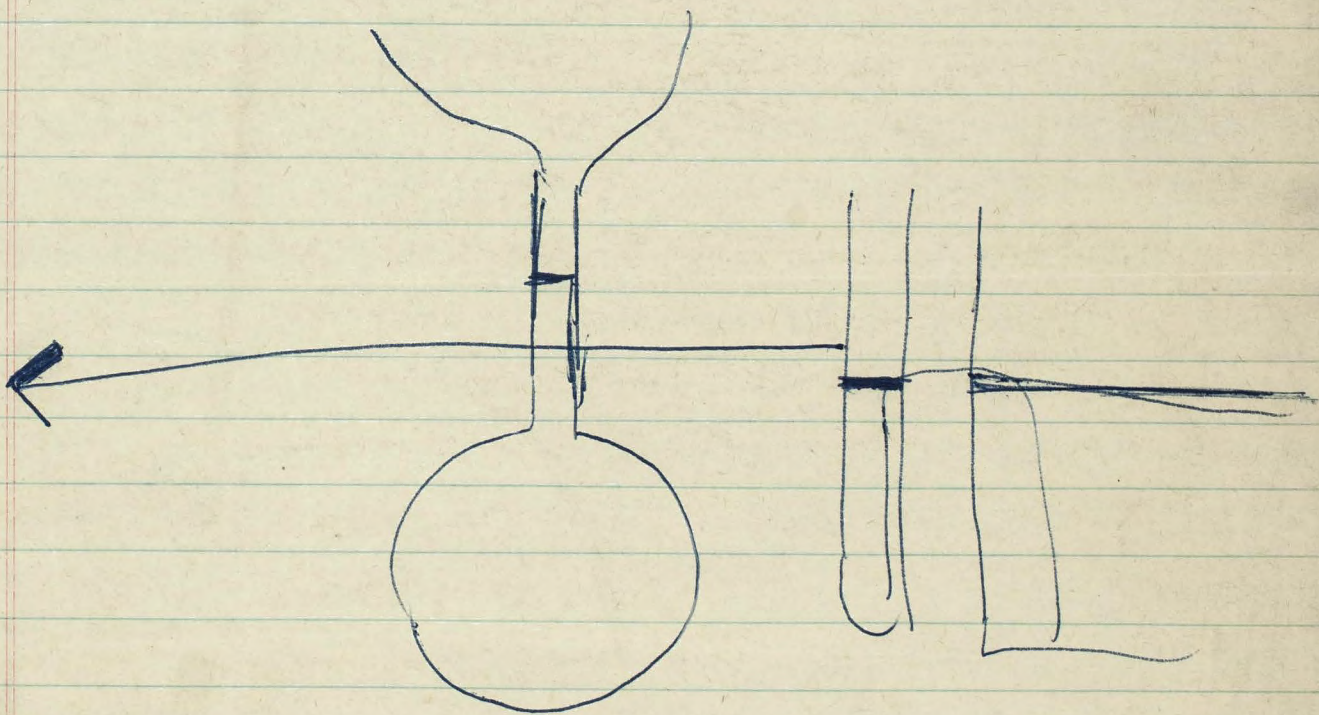
look in literature at U.V. irradiated  
population.

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Age population by U.V.  
irradiation

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1101 Belaire  
Apt 304



from Argonne Pl. 16 // 355 pm





Barrow

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Clemente Estable

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Houmas

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